

Chapter 13 States Of Matter Worksheet Answers

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Chapter 13: States of Matter Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist

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all matter consists of tiny particles that are constantly in motion What are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the particles in a gas are rapid, constant, and random.

Chapter 13: States of Matter - Chemistry by Anna

Chapter 13 States of Matter 137 SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

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Chemistry (12th Edition) answers to Chapter 13 - States of Matter - 13.1 The Nature of Gases - 13.1 Lesson Check - Page 424 8 including work step by step written by community members like you. Textbook Authors: Wilbraham, ISBN-10: 0132525763, ISBN-13: 978-0-13252-576-3, Publisher: Prentice Hall

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There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

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Chapter 13 Vocabulary - States of Matter. kinetic energy. kinetic theory. gas pressure. vacuum. the energy an object has due to its motion. all matter consists of tiny particles that are in constant mot.... results from the force exerted by a gas per unit surface area.... an empty space with no particles and no pressure.

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Chapter 13 States Of Matter Chemistry Test Answers

Chapter 13 - States of Matter - 13 Assessment - Page 443: 54 Answer Temperature stays constant because the energy is needed to break the intermolecular forces and changing the phase.

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Chapter 13 - States of Matter - 13 Assessment - Page 444: 58 Answer It is called a dynamic equilibrium because vaporization and condensation are still occurring at equal rates even though there is no net change in the number of particles in the liquid or vapor.

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Worksheet Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Chapter 13 States Of Matter Worksheet Look at the text on page 315 for the answer.

Chemistry (12th Edition) Chapter 13 - States of Matter ...

Chapter 13 States of Matter. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. bill_brossman. Terms in this set (26) Kinetic Energy. The energy an object has because of its motion. Kinetic Theory. A theory that explains the states of matter, based on the concept that all matter consists of tiny particles that are ...

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