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NaH_2PO_4 and Na_2HPO_4 in water, in which NaH_2PO_4 produces the acidic ion and Na_2HPO_4 produces the conjugate base ion. (i) Deduce the acid and conjugate base ions that make up the phosphate buffer and state the ionic equation that represents the phosphate buffer. [3] (ii) Describe how the phosphate buffer minimizes the effect of the addition of a

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- 7 - Turn over 4. A buffer solution with a pH of 3.87 contains 7.41 g dm⁻³ of propanoic acid, CH₃CH₂COOH, together with an unknown quantity of sodium propanoate, CH₃CH₂COONa. (a) Define the term buffer solution. [2]

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